**SUPPORTING INFORMATION**

Important role of native -cyclodextrin in the stabilization of transition metal salts.

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## Manganese

Figure 1S: UV-Vis spectrum of MnBr2 in NaOH solution at different time after preparation. [MnBr2 ] 1.6x10-4 M, [NaOH] 1.01x10-2 M

Figure 2S: UV-Vis spectrum of MnBr2 in NaOH solution in the presence of β-CD at different time after preparation. [MnBr2 ] 1.6x10-4 M, [-CD] 1.8x10-4 M [NaOH] 1.01x10-2 M



Figure 3S: UV-Vis spectrum of MnCl2 in NaOH solution at different time after preparation. [MnCl2] 1.7x10-4 M, [NaOH] 1.0x10-2 M.



Figure 4S: UV-Vis spectrum of MnCl2 in NaOH solution in the presence of β-CD at different time after preparation. [MnCl2] 1.8x10-4 M, [-CD] 1.7x10-4 M, [NaOH] 1.0x10-2 M.



Figure 5S: UV-Vis spectrum of Mn(NO3)2 in NaOH solution at different time after preparation. [Mn(NO3)2] 8.9x10-4 M, [NaOH] 9.9x10-3 M.



*Iron*

Figure 6S: UV-Vis spectrum of FeBr2 in NaOH solution at different time after preparation. [FeBr2] 1.0x10-3 M, [NaOH] 9.9x10-3 M.



Figure 7S: UV-Vis spectrum of FeCl3 in NaOH solution at different time after preparation. [FeCl3] 1.5x10-4 M, [NaOH] 1.0x10-2 M.

Figure 8S: UV-Vis spectrum of FeCl3 in NaOH solution in the presence of β-CD at different time after preparation. [FeCl3] 2.8x10-4 M, [-CD] 1.7x10-4 M, [NaOH] 1.0x10-2 M.



Figure 9S: UV-Vis spectrum of Fe(NO3)3•9H2O in NaOH solution at different time after preparation. [Fe(NO3)3·9H2O] 1.4x10-4 M, [NaOH] 1.1x10-2M.



Figure 10S: UV-Vis spectrum of Fe(NO3)3•9H2O in NaOH solution in the presence of β-CD at different time after preparation. [Fe(NO3)3·9H2O] 1.3x10-4 M, [-CD] 1.7x10-4 M, [NaOH] 1.0x10-2 M.



## Cobalt

Figure 11S: UV-Vis spectrum of CoBr2 in NaOH solution at different time after preparation. [CoBr2 ] 1.3x10-4 M, [NaOH] 1.01x10-2 M.



Figure 12S: UV-Vis spectrum of CoBr2 in NaOH solution in the presence of β-CD at different time after preparation. [CoBr2] 2.1x10-3 M, [-CD] 1.0x10-2 M, [NaOH] 1.0x10-2 M. [-CD]/[CoBr2] = 4.8 ratio. The longer arrow indicates the increase in absorbance observed at the beginning of the study.



Figure 13S: UV-Vis spectrum of CoCl2 in 10-2 M NaOH solution at different time after preparation. [CoCl2]1.2x10-3 M, [NaOH] 9.9x10-3 M.



Figure 14S: UV-Vis spectrum of CoCl2 in NaOH solution in the presence of β-CD at different time after preparation. [CoCl2] 1.0x10-3 M, [-CD] 1.0x10-2 M, [NaOH] 9.9x10-3 M. [-CD]/[CoCl2] =10 ratio. The arrow indicates the increase in absorbance observed at the beginning of the study.



Figure 15S: UV-Vis spectrum of Co(NO3)2·6H2O in NaOH solution at different time after preparation. [Co(NO3)2·6H2O] 1.7x10-4 M, [NaOH] 1.0x10-2 M.



Figure 16S: UV-Vis spectrum of Co(NO3)2·6H2O in 10-2 M NaOH solution in the presence of β-CD at different time after preparation. [Co(NO3)2·6H2O] 1.5x10-4 M, [-CD] 1.7x10-4 M, [NaOH] 1.0x10-2 M.

*Nickel*



Figure 17S: UV-Vis spectrum of NiBr2 in NaOH solution in the presence of β-CD at different time after preparation. [NiBr2] 1.9x10-4 M, [-CD] 1.8x10-4 M, [NaOH] 1.01x10-2 M.



Figure 18S: UV-Vis spectrum of NiCl2 in NaOH solution at different time after preparation. [NiCl2] 1.0x10-3 M, [NaOH] 9.9x10-3 M.



Figure 19S: UV-Vis spectrum of NiCl2 in NaOH solution in the presence of β-CD at different time after preparation. [NiCl2] 9.9x10-4 M, [-CD] 1.0x10-3 M, [NaOH] 9.9x10-3 M.



Figure 20S: UV-Vis spectrum of NiCl2 in NaOH solution in the presence of β-CD at different time after preparation. [NiCl2] 9.9x10-4 M, [-CD] 1.0x10-2 M, [NaOH] 9.9x10-3 M. [-CD]/[NiCl2] = 10.1 ratio



Figure 21S: UV-Vis spectrum of NiCl2 in NaOH solution in the presence of β-CD at different time after preparation. [NiCl2] 4.9x10-4 M, [-CD] 1.0x10-2 M, [NaOH] 9.9x10-3 M. [-CD]/[NiCl2] = 20.4 ratio



Figure 22S: UV-Vis spectrum of Ni(NO3)2·6H2O in NaOH solution at different time after preparation. [Ni(NO3)2·6H2O] 1.9x10-4 M, [NaOH] 1.0x10-2 M.



Figure 23S: UV-Vis spectrum of Ni(NO3)2·6H2O in NaOH solution and in the presence of -CD at different time after preparation. [Ni(NO3)2·6H2O] 1.8x10-4 M, [-CD] 1.7x10-4 M, [NaOH] 1.0x10-2 M.



Figure 24S: UV-Vis spectrum of Ni(NO3)2·6H2O in NaOH solution and in the presence of -CD at different time after preparation. [Ni(NO3)2·6H2O] 2.0x10-3 M, [-CD] 1.0x10-2 M, [NaOH] 9.9x10-3 M. [-CD]/[ Ni(NO3)2] = 5 ratio



*Copper*

Figure 25S: UV-Vis spectrum of CuBr2 in NaOH solution at different time after preparation. [CuBr2] 2.1x10-4 M, [NaOH] 1.0x10-2 M.



Figure 26S: UV-Vis spectrum of CuBr2 in NaOH solution and in precense of -CD at different time after preparation. [CuBr2] 2.1x10-4 M, [-CD] 1.7x10-4 M, [NaOH] 1.0x10-2 M.

Figure 27S: UV-Vis spectrum of CuBr2 in NaOH solution and in precense of -CD at different time after preparation. [CuBr2] 1.0x10-4 M, [-CD] 1.0x10-3 M, [NaOH] 9.9x10-3 M.

[-CD]/[CuBr2] = 10 ratio



Figure 28S: UV-Vis spectrum of Cu(NO3)2·2.5H2O in 10-2 M NaOH solution at different time after preparation. [Cu(NO3)2·2.5H2O] 1.8x10-4 M, [NaOH] 1.0x10-2 M.



Figure 29S: UV-Vis spectrum of Cu(NO3)2·2.5H2O in NaOH solution in presence of -CD at different time after preparation. [Cu(NO3)2·2.5H2O] 1.6x10-4 M, [-CD] 1.7x10-4 M, [NaOH] 1.0x10-2 M.



*Zinc*

Figure 30S: UV-Vis spectrum of ZnBr2 in NaOH solution at different time after preparation. [ZnBr2 ] 2.0x10-4 M, [NaOH] 1.0x10-2 M.



Figure 31S: UV-Vis spectrum of ZnBr2 in NaOH solution and in presence of -CD at different time after preparation. [ZnBr2] 2.0x10-4 M, [-CD] 1.8x10-4 M, [NaOH] 1.0x10-2 M.



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